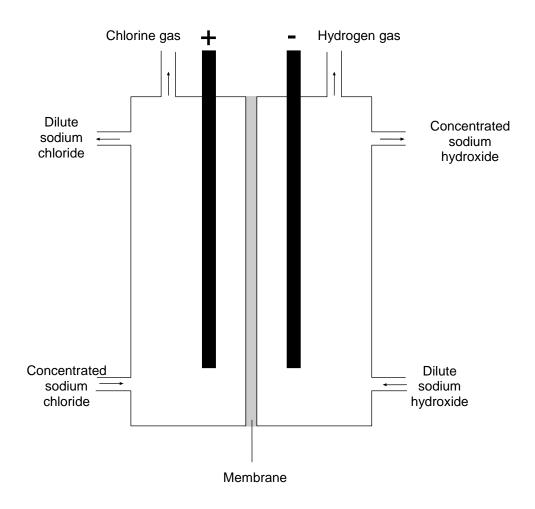
Chemicals from Salt: Answers



- 1. (a) 1 mark per correct label
 [7]

 (b) Sodium ions
 [1]

 (c) $2CI^- \rightarrow CI_2 + 2e^-$ [1]

 $2H_2O + 2e^- \rightarrow H_2 + 2OH^-$ [1]

 (d) Mercury is expensive
 [1]
- (a) One mole (formula mass) of hydrogen and one mole (formula mass) of chlorine produced [1]
 Chlorine atom 35.5 times heavier than hydrogen atom [1]
 (b) Two moles (formula masses) KCl produce 2 moles (formula masses) KOH [1]
 149 tonnes → 112 tonnes [1]
 Mass of potassium chloride = 149 x 5.6 = 7.45 tonnes [1]

112

	(c)	Named apparatus – burette, pipette	
		Measure out 25 cm ³ of potassium hydroxide solution	
		Add named indicator	
		Add hydrochloric acid in small portions	
		Until indicator changes colour	
		Repeat until consistent results	
		Mark any five of the above points	[5]
		$KOH + HCI \rightarrow KCI + H_2O$	[1]
3.	(a)	Hole drilled down to salt deposits	[1]
		Cold water pumped down to deposits	[1]
		Salt dissolves and salt solution is pumped to the surface	[1]
	(b)	Salt dissolves in water, coal does not	[1]
	(c)	Salt has almost the same solubility at all temperatures	[1]
		Extra costs of heating the water cannot be recovered	[1]
	(d)	Any possible problem – eg subsidence caused by salt being remo	oved underground [1]
4.	(a)	Salt (sodium chloride) and water or brine	[1]
		Sodium hydroxide, hydrogen and chlorine	[2]
		Two marks for three products and one mark for one or two	
	(b)	Salt is available close to the site	
		Brine can be delivered to the factory by pipeline	
		Close to other chemical factories which might use the products	
		Specific mention of use of one of the products	
		Ports close to factory for exporting products	
		Other good distribution systems - eg road, rail and canal	
		Suitable labour force in the area	
		Mark any five of the above points or alternatives	[5]

