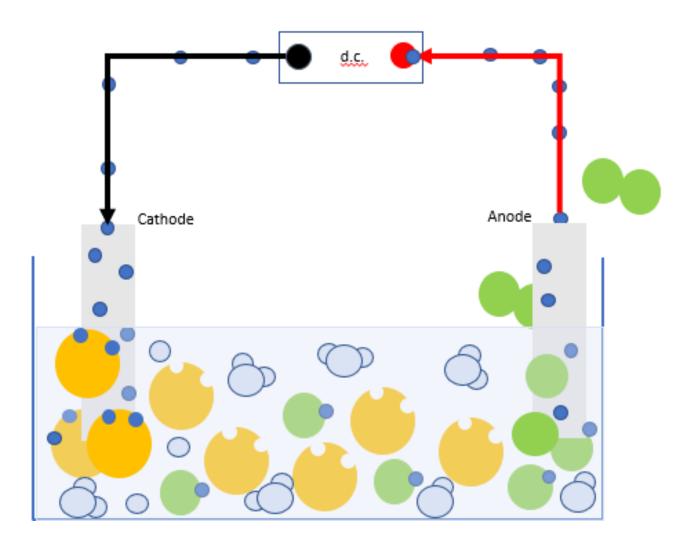
Need help? Use these prompts









$$H_2O$$
 \rightleftharpoons $H^+(ag) + HO^-(ag)$

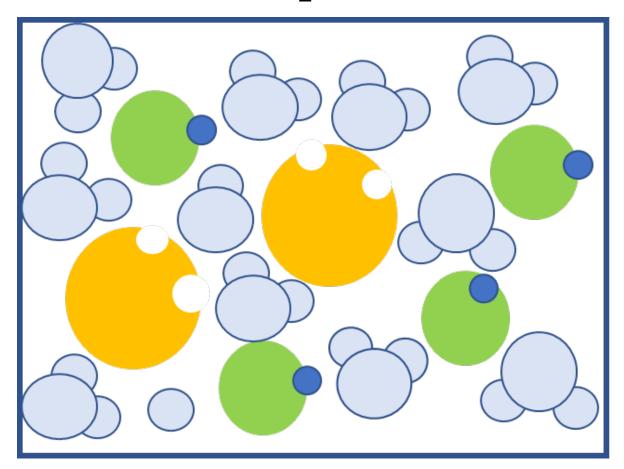
Cathode
$$Cu^{2+}(ag) + 2e^{-} \longrightarrow Cu(s)$$

Anode 2
$$Cl_2(g) + 2e^-$$





copper(II) chloride solution CuCl₂(aq)











(g) (aq)
$$\longrightarrow$$
 +

$$H_2O(I)$$
 \longrightarrow $H^+ (aq) + HO^-(aq)$

$$Cu^{2+}(aq) + 2e^{-} \longrightarrow Cu(s)$$

$$2 \text{ Cl}^{-}(aq) \rightarrow \text{Cl}_{2}(g) + 2e^{-}$$

2232-4				
electrolyte	graphite	molecule	inference	wire
anode	electrode	electron	gain	positive
cathode	electrolysis	aqueous solution	loss	negative
inert	atom	external circuit	move toward	blue litmus
half equation	ion	ionic equation	attract	bleach
oxidation	reduction	observation	terminal	copper(II) chloride