# Exam reflection sheet

***Education in Chemistry***September 2019
rsc.li/2X9cDvy

Simply giving a student a grade after an exam doesn’t tell them how to improve. Completing a reflection sheet gives them the opportunity to revisit and consider their approach to revision and exam answers. Here is an example sheet and a blank one for you to adapt.

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| **Name**: |  |

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| **How confident did you feel before you sat the chemistry paper 1 exam?**  |
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| **Did you revise? What techniques did you find useful?** |
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| **Chemistry paper 1H** |
| **Q** | **Topic(s)** | **RG page(s)** | **Tasks** |
| 1 | Halogens | 25 | Explain why the halogens get less reactive as you go down the group. |
| 2 | Transition metals and alkali metals | 24 | Compare the properties of transition metals with the alkali metals. |
| 3 | Ionic bonding and properties of ionic compoundsCovalent bonding and simple molecular substances | 28-3031-32 | 1. Explain why ionic compounds conduct electricity when molten or aqueous but not when solid.
2. Explain why ionic compounds have high melting points.
3. Draw a dot cross diagram for N2.
4. Explain why N2 has a low boiling point.
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| 4 | Distillation | 18 | Explain how distillation works.  |
| 5 | Making salts | 54 | Describe how a sample of pure, dry potassium nitrate crystals could be made. Watch the Making salts required practical video and use the pages in the revision guide for support. |
| 6 | Cells, batteries and hydrogen fuel cells | 64-65 | Complete questions 11-20 on page 66. |
| 7 | Rates of reaction | 67-71 | Watch the Rates of reaction required practical videos. Write a method for each. List the independent, dependent and control variables and any sources of error.  |
| 8 | Titration calculations | 52-53 | Complete questions 4-7 on page 60. |
| 9 | Bond energies | 63 | Complete questions 9-10 on page 66. |

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| **Q** | **Topic(s)** | **RG Page(s)** | **Tasks** |
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