

## Exam reflection sheet

*Education in Chemistry*

September 2019

rsc.li/2X9cDvy



**Simply giving a student a grade after an exam doesn't tell them how to improve. Completing a reflection sheet gives them the opportunity to revisit and consider their approach to revision and exam answers. Here is an example sheet and a blank one for you to adapt.**

<b>Name:</b>	
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<b>How confident did you feel before you sat the chemistry paper 1 exam?</b>
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<b>Did you revise? What techniques did you find useful?</b>
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<b>Chemistry paper 1H</b>			
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<b>Q</b>	<b>Topic(s)</b>	<b>RG page(s)</b>	<b>Tasks</b>
1	Halogens	25	Explain why the halogens get less reactive as you go down the group.
2	Transition metals and alkali metals	24	Compare the properties of transition metals with the alkali metals.
3	Ionic bonding and properties of ionic compounds  Covalent bonding and simple molecular substances	28-30  31-32	<ol style="list-style-type: none"> <li>1. Explain why ionic compounds conduct electricity when molten or aqueous but not when solid.</li> <li>2. Explain why ionic compounds have high melting points.</li> <li>3. Draw a dot cross diagram for N<sub>2</sub>.</li> <li>4. Explain why N<sub>2</sub> has a low boiling point.</li> </ol>
4	Distillation	18	Explain how distillation works.
5	Making salts	54	Describe how a sample of pure, dry potassium nitrate crystals could be made. Watch the Making salts required practical video and use the pages in the revision guide for support.
6	Cells, batteries and hydrogen fuel cells	64-65	Complete questions 11-20 on page 66.
7	Rates of reaction	67-71	Watch the Rates of reaction required practical videos. Write a method for each. List the independent, dependent and control variables and any sources of error.
8	Titration calculations	52-53	Complete questions 4-7 on page 60.
9	Bond energies	63	Complete questions 9-10 on page 66.

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