## **Spacing grid example**

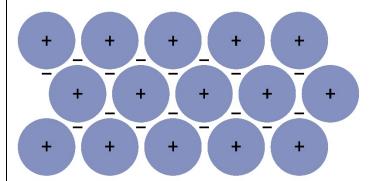


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Use a spacing grid like this to give pupils retrieval cues.

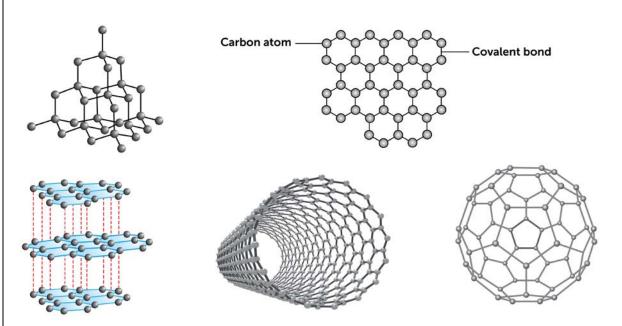
Draw the bonding in sodium chloride (NaCl) and water (H <sub>2</sub> O).			
Draw the boliding in Sodidin Chloride (NaCi) and water (1120).			
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Na Na Sodium	35.5 <b>Cl</b> Chlorine 17	1 <b>H</b> Hydrogen 1	Oxygen 8
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Explain how sodium chloride has a high melting and boiling point, but water has a low melting and boiling point.			
What other properties do ionic compounds have?			

Using this model of closely packed positive metal ions and a sea of delocalised electrons, explain why metals:



- Have high melting and boiling points:
- Are good conductors of heat and electricity:
- Are malleable and ductile:

Allotropes of carbon – name each one. What can you remember?



How are the above allotropes (other than diamond) able to conduct electricity?