# Structure strips for rates of reaction

***Education in Chemistry***September 2019
[rsc.li/2P0JDlW](https://rsc.li/2P0JDlW)

Scaffolding helps students overcome the fear of a blank page. Structure strips provide suitable prompts for a piece of writing, in this case rates of reaction. The student sticks the strip into the margin of their exercise book and writes alongside it.

See the next page.

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| --- | --- | --- | --- | --- |
| **Rates structure strip***Define ‘rate of reaction’.* | **Rates structure strip***Define ‘rate of reaction’.* | **Rates structure strip***Define ‘rate of reaction’.* | **Rates structure strip***Define ‘rate of reaction’.* | **Rates structure strip***Define ‘rate of reaction’.* |
| *Give four factors that affect the rate of reaction.* | *Give four factors that affect the rate of reaction.* | *Give four factors that affect the rate of reaction.* | *Give four factors that affect the rate of reaction.* | *Give four factors that affect the rate of reaction.* |
| *Name the two general ways of measuring rate of reaction.* | *Name the two general ways of measuring rate of reaction.* | *Name the two general ways of measuring rate of reaction.* | *Name the two general ways of measuring rate of reaction.* | *Name the two general ways of measuring rate of reaction.* |
| *Explain how collision theory explains rate of reaction.* | *Explain how collision theory explains rate of reaction.* | *Explain how collision theory explains rate of reaction.* | *Explain how collision theory explains rate of reaction.* | *Explain how collision theory explains rate of reaction.* |
| *Describe how the following rates of reaction could be determined (include equations and diagrams where necessary).**(i) Mg + HCl**(ii) Na2S2O3 + HCl* | *Describe how the following rates of reaction could be determined (include equations and diagrams where necessary)**(i) Mg + HCl**(ii) Na2S2O3 + HCl* | *Describe how the following rates of reaction could be determined (include equations and diagrams where necessary)**(i) Mg + HCl**(ii) Na2S2O3 + HCl* | *Describe how the following rates of reaction could be determined (include equations and diagrams where necessary)**(i) Mg + HCl**(ii) Na2S2O3 + HCl* | *Describe how the following rates of reaction could be determined (include equations and diagrams where necessary)**(i) Mg + HCl**(ii) Na2S2O3 + HCl* |
| *Define the term catalyst.* | *Define the term catalyst.* | *Define the term catalyst.* | *Define the term catalyst.* | *Define the term catalyst.* |
| *Give some examples of catalysts and the reactions they catalyse.* | *Give some examples of catalysts and the reactions they catalyse.* | *Give some examples of catalysts and the reactions they catalyse.* | *Give some examples of catalysts and the reactions they catalyse.* | *Give some examples of catalysts and the reactions they catalyse.* |