

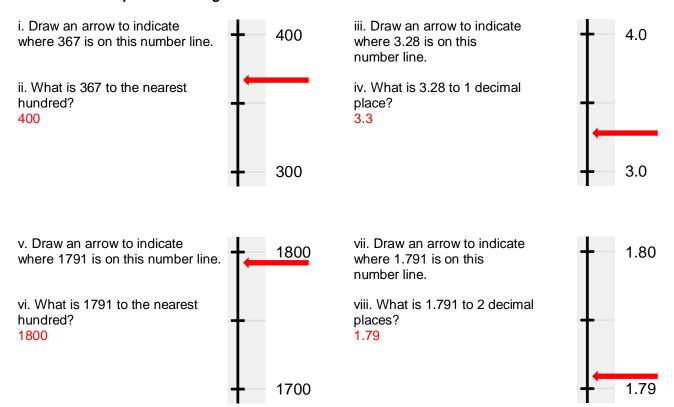
Expressing uncertainty in data

Education in Chemistry July 2020 rsc.li/2Xfd61c

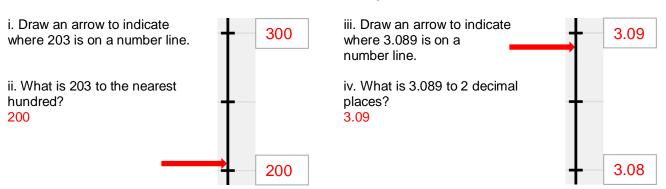
Practise handling significant figures, decimal places and uncertainty in measurements.

1. Rounding

a. Answer the questions using the number lines.



b. Fill in the blanks on the number lines and answer the questions.



c. Circle the number that is the nearest 1000, 100 and 10 to the number mentioned in each question and indicate whether it is rounded up or down.

i. 2527

up or down?

nearest	nearest	nearest
1000	100	10
(3000)	2600	2530
2000	2500	2520
up	down	up

ii. 4458

	nearest	nearest	nearest
	1000	100	10
	5000	4500	(4460)
	4000	4400	4450
up or down?	down	up	up

d. Circle the number that is the nearest integer, 0.1 and 0.01 and indicate whether it is rounded up or down.

iii. 2.527

	nearest	nearest	nearest
	integer	0.1	0.01
	(3)	2.6	2.53
	2	2.5	2.52
up or	up	down	up

iv. 4.458

	nearest	nearest	nearest
	integer	0.1	0.01
	5	4.5	4.46
	(4)	4.4	4.45
up or down?	down	up	up

e. Insert the missing numbers so that the table looks like the ones above. Then, circle the number that is the nearest integer, 0.1 and 0.01 and indicate whether it is rounded up or down.

i. 4.097

	nearest integer	nearest 0.1	nearest 0.01
	5	4.1)	(4.10)
	4	4.0	4.09
up or down?	down	up	up

ii. 10.893

	nearest	nearest	nearest
	integer	0.1	0.01
	(11)	(10.9)	10.90
	10	10.8	(10.89)
up or down?	up	up	down

2. Significant figures

Circle the correct answer.

a. Round 34.59 to 1 significant figure					
A: 34	B: 30	C: 35	D: 3	E: 34.5	
b. Round 35,683	to 1 significant figure)			
A: 35683	B: 30000	C: 40000	D: 3	E: 4	
-					
c. Round 76.984 to	o 3 significant figures	3			
A: 77.0	B: 76.9	C: 76.984	D: 77	E: 76	
d. Round 0.003865 to 1 significant figure					
A: 0	B: 0.003	C: 0.004	D: 0.00387	E: 1	
f. Round 0.003865 to 3 significant figures					
A: 0.00	B: 0.003	C: 0.004	D: 0.00386	E: 0.00387	

3. Chemistry contexts

- a. Calculate the rate of reaction that produces 25.0 cm³ of gas in 5.90 s. $25.0 \div 5.90 = 4.24$ cm³ s⁻¹
- b. Calculate the rate of reaction that produces 25 cm³ of gas in 5.90 s. $25 \div 5.90 = 4.2 \text{ cm}^3 \text{ s}^{-1}$
- c. Calculate the rate of reaction that produces 25.0 cm³ of gas in 5.9 s. $25.0 \div 5.9 = 4.2 \text{ cm}^3 \text{ s}^{-1}$
- d. Calculate the rate of reaction that produces 25 cm³ of gas in 5.9 s. $25 \div 5.9 = 4.2$ cm³ s⁻¹
- e. Calculate the concentration of a sodium chloride solution when 9.80 g of solid was dissolved in 100 cm³ water.

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9.80 \div 100 = 0.1 \text{ g cm}^{-3}
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f. Calculate the concentration of a sodium chloride solution when 9.8 g of solid was dissolved in 100 cm³ water.

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9.8 \div 100 = 0.1 \text{ g cm}^{-3}
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g. Calculate the concentration of a sodium chloride solution when 0.98 g of solid was dissolved in 10 cm³ water.

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0.98 \div 10 = 0.1 \text{ g cm}^{-3}
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Calculate the concentration of a sodium chloride solution 0.98 g of solid was dissolved in 10.0 cm³ water

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0.98 \div 10.0 = 0.098 \text{ g cm}^{-3}
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