

Equilibrium reactions Worksheet

Read each of the statements below.

Tick one of the boxes to show that you think the statement is **true**, **false** or that you are **unsure** whether it is true or not.

1. When a reactant is added to a system in equilibrium, the forward reaction will occur to use up all the added material and so restore the equilibrium.
2. When a reactant is added to a system in equilibrium, more product is produced but the value of the equilibrium constant, K , remains unaltered.
3. A system reaches equilibrium when the concentration of the reactants is equal to the concentration of the products.
4. A high value of the equilibrium constant, K , means that the forward reaction is very fast.
5. When the reactants of an equilibrium system are mixed together the rate of the forward reaction increases until equilibrium is established.
6. A catalyst increases the rate of reaction of both the forward and reverse reactions.
7. In an equilibrium system in which the forward reaction is endothermic, the reverse reaction is exothermic.
8. If a system in equilibrium where the forward reaction is endothermic is heated, the rate of the forward reaction increases but the rate of the reverse reaction decreases.

Statement	The statement is true	The statement is false	I am unsure
1			
2			
3			
4			
5			
6			
7			
8			

For each statement that you think is false, explain why you think it is wrong.

Write down any changes in your thinking as a result your discussions.