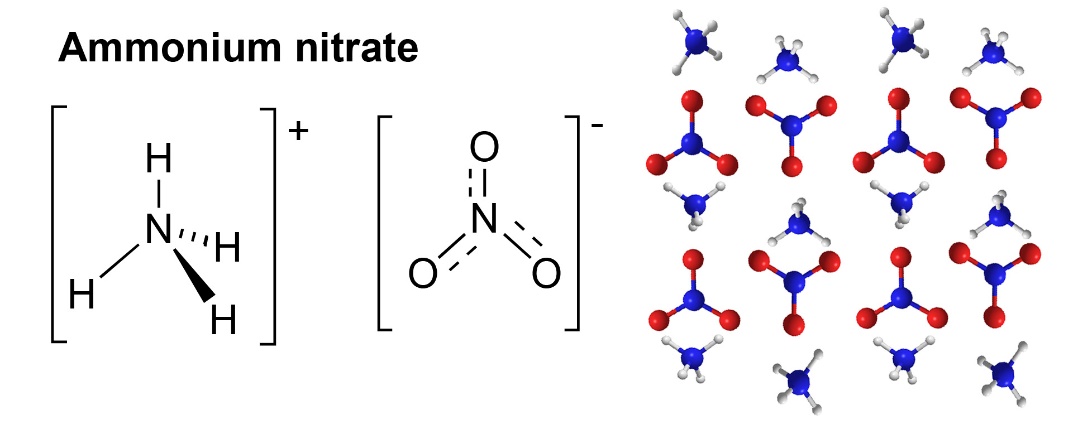
# Ammonium nitrate explosions

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[rsc.li/2TdpNqC](https://rsc.li/2TdpNqC)

The decomposition of ammonium nitrate at temperatures above 260oC leads to a ‘runaway’ explosive reaction. Many aspects of this ‘runaway’ reaction can be explained by concepts such as enthalpy change and by taking a closer look at the structure and bonding of ammonium nitrate.

## Task 1 – Structure and bonding in ammonium nitrate

The structure and geometry of ammonium nitrate are shown in the diagrams below.

1. Ammonium nitrate is composed of two polyatomic ions – identify the different type(s) of bonding within its structure.
2. Using the VSEPR theory and the diagrams above, state and explain the geometries of the NH4+ and NO3- ions shown.

## Task 2 – Decomposition or explosion?

Pure ammonium nitrate does not explode easily and can be handled safely. It decomposes at 230oC producing nitrous oxide gas (N2O) and water vapour.

1. Give a balanced equation for the decomposition reaction.

Above 260oC, if confined and when contaminated, ammonium nitrate will explode forming toxic gases such as NO2, responsible for the ‘orange brown fireball’ described in the article.

The following equation represents one of the reactions contributing to the explosion**:**

4NH4NO3(s)→ 3N2(g) + 2NO2(g) + 8H2O(g)

1. Given that the molar gas volume is 24.5 dm3 at 298K and 1atm calculate the total volume of gas produced from 80 kg of NH4NO3(s) under these conditions.
2. Use your value to calculate the total volume of gas at the same pressure and 300oC.  
   Scale this up to calculate the vast volume of gas produced from 2750 tonnes of NH4NO3(s) in the Beirut explosion.

## Task 3 – Calculating the enthalpy of reaction, ΔHR, and using kinetic theory to explain why this is a ‘runaway’ reaction

The data below will be used in Tasks 3 - 5

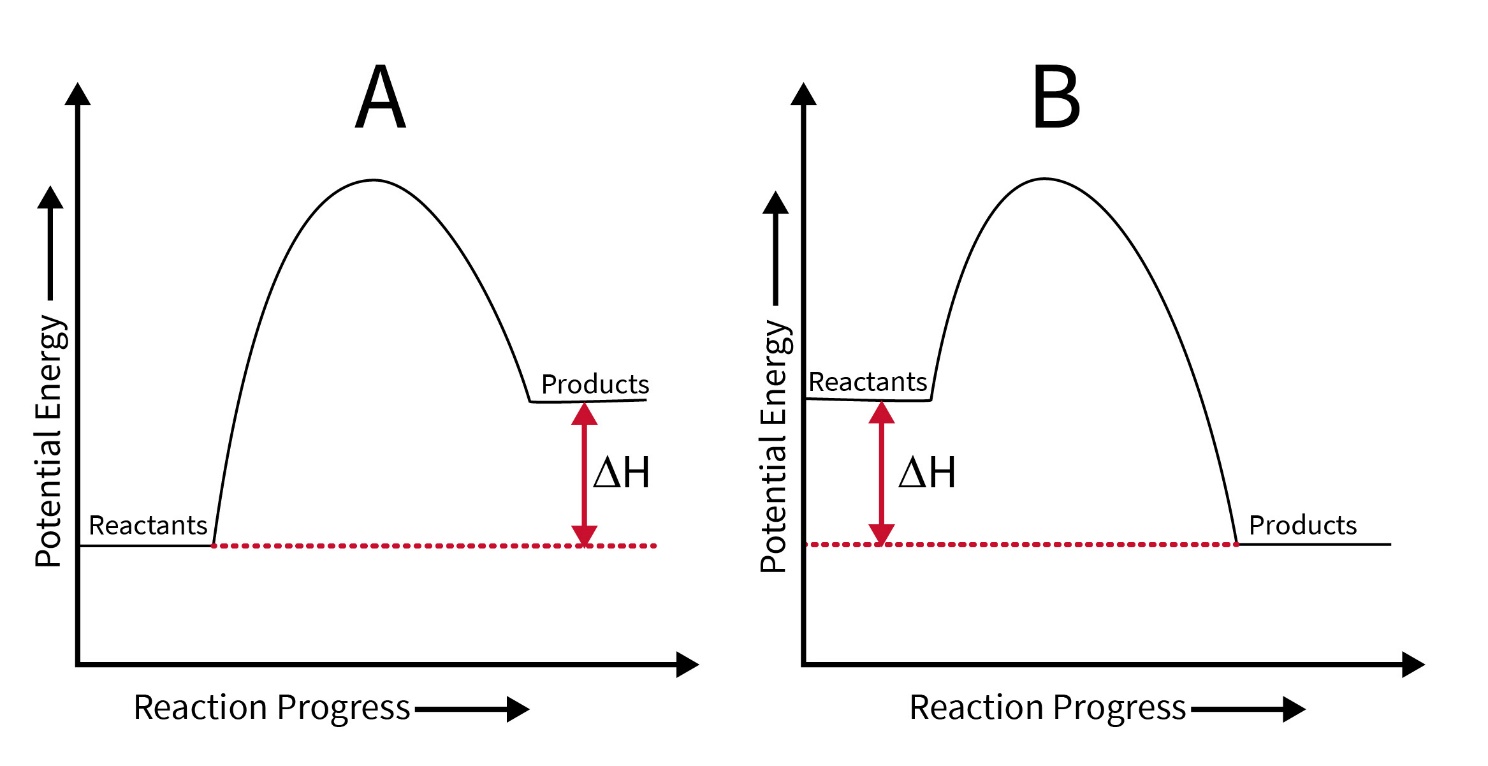
|  |  |  |
| --- | --- | --- |
| **Compound** | **ΔH0f / KJmol-1** | **S0/JK-1mol-1** |
| NH4NO3 (s) | -365.6 | +15.1 |
| N2(g) | 0 | +153.3 |
| NO2(g) | +33.2 | +240.1 |
| H2O(g) | -241.8 | +188.8 |

Once decomposition begins a ‘runaway’ reaction occurs. In a runaway reaction an exothermic reaction goes out of control. The heat evolved raises the temperature of the reacting mixture leading to an increase in reaction rate, which causes a further increase in temperature and a further increase in reaction rate until an explosion occurs.

1. Use the standard enthalpies of formation, ΔH0f**,** given in KJmol-1 to calculate the enthalpy change for the reaction and confirm that it is exothermic

4NH4NO3(s)→ 3N2(g) + 2NO2(g) + 8H2O(g)

1. Select which of the energy profile templates below (A or B) correctly represents the enthalpy change you have calculated and label it showing reactants, products, ΔHR and Eact.



1. Use kinetic theory to explain why an increase in temperature causes an increase in reaction rate.

## Task 4 – Calculating the entropy change, ΔS, for the reaction

The changes in state during the decomposition reaction below suggest there will be a significant increase in the disorder**,** entropy, for the reaction – ie ΔS will be highly positive.

1. Use the standard entropies, S**0,** g**i**ven in JK-1mol-1 to confirm this is the case.

4NH4NO3(s)→ 3N2(g) + 2NO2(g) + 8H2O(g)

## Task 5 – Assessing the spontaneity of the reaction

To assess the spontaneity of a reaction both enthalpy and entropy changes need to be considered together in Gibb’s Equation.

Using your enthalpy and entropy changes from Tasks 3 and 4 and Gibbs equation below:

1. Comment on whether the decomposition reaction is likely to be spontaneous at all temperatures.
2. Using temperatures of 300**0**C and 5000C in amodel calculation comment on whether the decomposition is likely to become more favourable as the temperature increases.

### Reminders about Gibbs equation:

**Δ**G = **Δ**H -T**Δ**S

For a reaction to be spontaneous ΔG must be negative.  
Take care with units!  
ΔG and ΔH in KJmol-1, ΔS in JK-1mol-1 and T is in K.