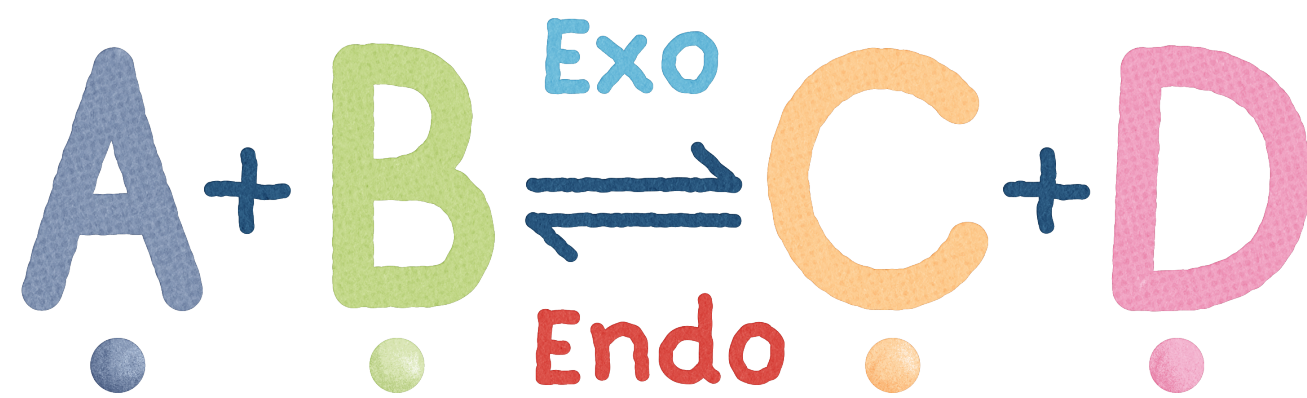
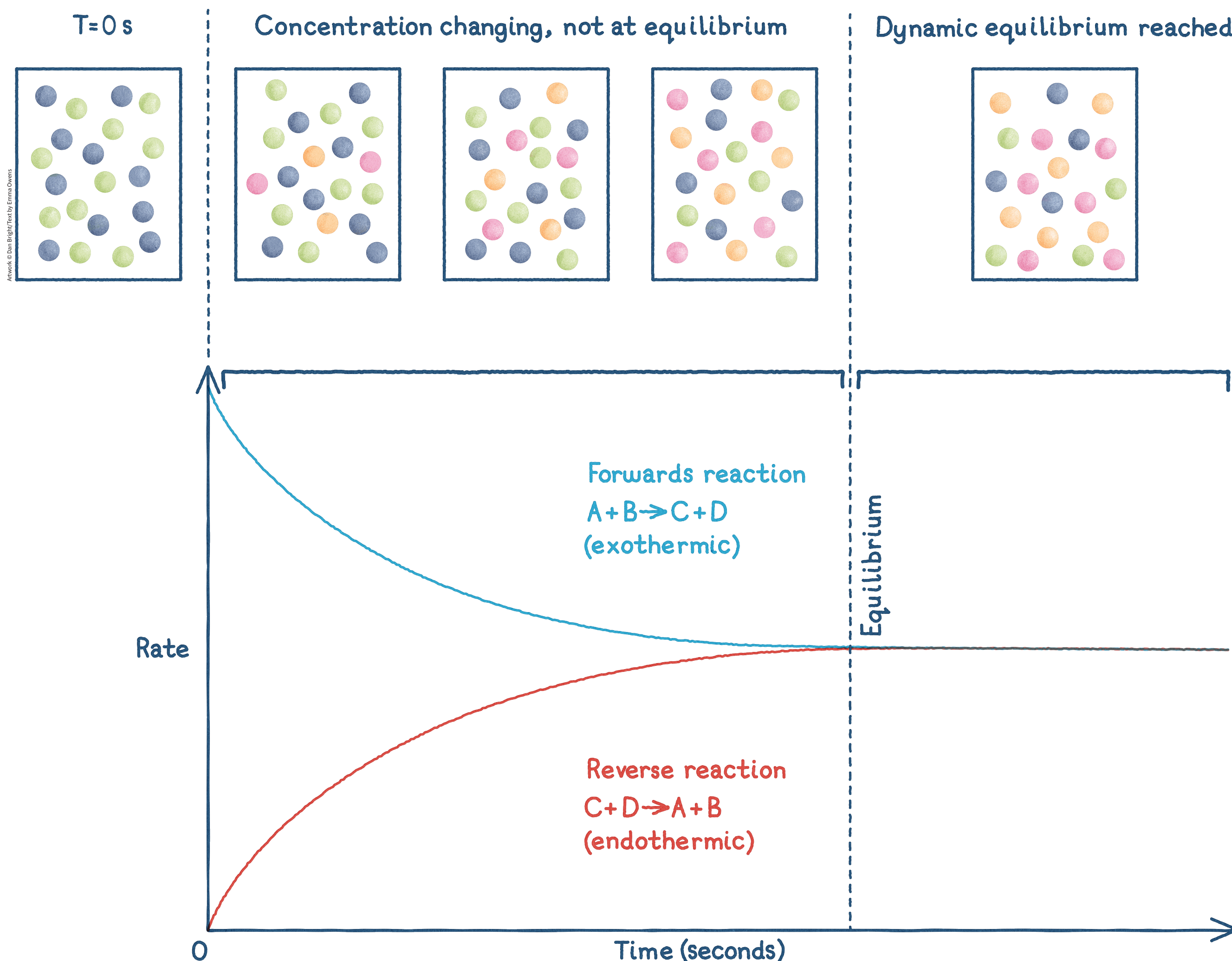


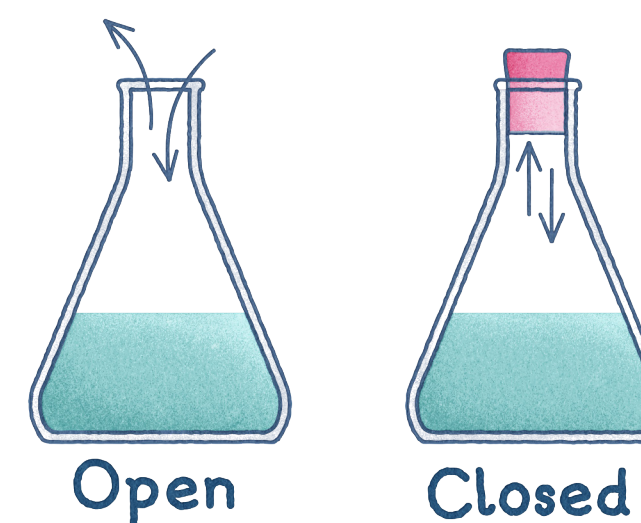
Chemical equilibrium



Equilibrium is an important process in industry, eg for making ammonia using the Haber process which is a **reversible reaction**. To make reversible reactions as efficient and sustainable as possible, manufacturers need to understand **equilibrium**. Because the **equilibrium position** – the **concentrations** of substances present at equilibrium – affects the **yield** of the product.



Note
Concentration of reactants (A+B) and products (C+D) do not have to be equal at equilibrium.



Dynamic equilibrium can only be reached in a **closed system**. For example, a reaction flask where nothing can enter or leave.