Chemical equilibrium

Equilibrium is an important process in industry, e.g. for making ammonia using the Haber process which is a reversible reaction. To make reversible reactions as efficient and sustainable as possible, manufacturers need to understand equilibrium. Because the equilibrium position – the concentrations of substances present at equilibrium – affects the yield of the product.

Note
Concentration of reactants (A+B) and products (C+D) do not have to be equal at equilibrium.

Dynamic equilibrium can only be reached in a closed system. For example, a reaction flask where nothing can enter or leave.