**Cracking: knowledge check**

1. The data in the table compares the supply of crude oil fractions from fractional distillation with their demand.

|  |  |  |
| --- | --- | --- |
| **Fraction** | **Percentage supplied from crude oil**  **()** | **Percentage demand**  **()** |
| gases | 3 | 6 |
| petrol | 10 | 27 |
| kerosene | 11 | 19 |
| bitumen | 12 | 4 |
| diesel | 15 | 23 |
| fuel oil | 49 | 21 |

Use the data in the table, and the words provided, to complete the sentences.

**gases petrol kerosene diesel fuel oil bitumen**

* 1. The fractions where the amount needed (demand) is higher than the amount produced (supply) are \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ , \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ , \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ and \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.
  2. The fractions where demand is higher than supply are \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ and \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.

1. Decide whether each of the statements is true or false and write your answer in the box provided.
   1. The process of cracking is used to break larger hydrocarbon molecules into smaller ones. **True False**
   2. Smaller hydrocarbon molecules are generally less useful than larger hydrocarbon molecules. **True False**
   3. Saturated hydrocarbons contain at least one double covalent bond between carbon atoms. **True False**
   4. Alkenes are saturated compounds. **True False**
   5. Alkane molecules only contain single covalent bonds between carbon atoms. **True False**
   6. Alkanes and alkenes are part of the same homologous series.

**True False**

1. Choose the correct words to complete the sentences.

The process of cracking is used to help match the supply of crude oil fractions with their **demand**/**price**.

The smaller, saturated **alkanes**/**alkenes** produced in cracking are used as **fuels**/**polymers** such as petrol, kerosene and **bitumen**/**diesel**.

The unsaturated **alkanes**/**alkenes** produced are used in the petrochemical industry to make other products including medicines, dyes and **fuels**/**polymers**.

1. Use the words provided to complete the sentences describing the two types of cracking.

**steam covalent speed up 800°C**

**550°C catalyst smaller**

During the process of cracking some of the \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ bonds in the molecules are broken to produce \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ alkanes and alkenes.

Catalytic cracking involves heating the hydrocarbons to a temperature of \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_to vaporise them. The vapours are passed over a \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ of zeolite to \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ the reaction.

Steam cracking involves heating the hydrocarbons to a temperature of \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ to vaporise them.

The vaporised hydrocarbons are mixed with \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ and cracking occurs.

**Cracking: test myself**

**Answer questions 2.1 to 2.7 by circling the correct answer.**

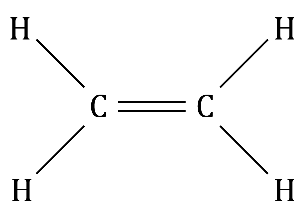
1. What type of reaction is cracking?

**reduction oxidation precipitation thermal decomposition**

1. What is the general molecular formula of an alkene?

**C*n*H2*n* C*n*H2*n* + 2 C*n*H2*n*O*n* CH**

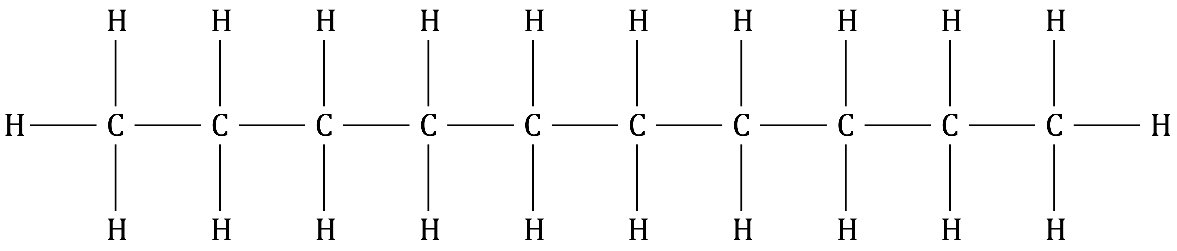
1. The image shows the displayed structural formula of ethene.



How many pairs of electrons are shared in the double covalent bond?

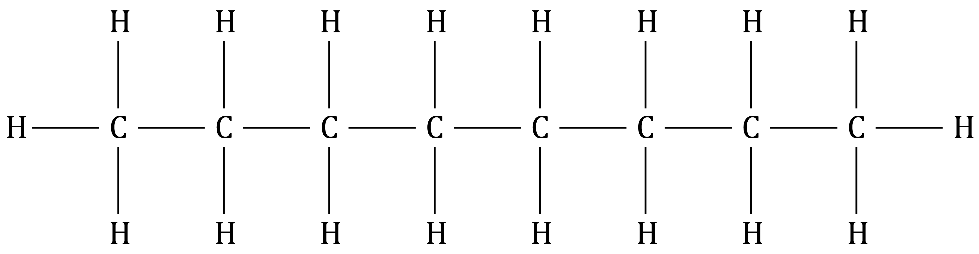
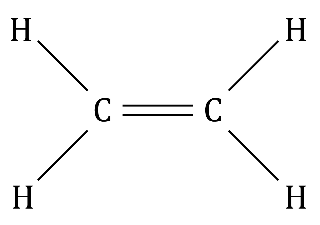
**one two three four**

1. This equation represents the cracking of decane, C10H22.



C10H22  decane

**Heat**



+

C8H18  octane C2H4  ethene

What is the unsaturated product?

**decane octane ethane ethene**

1. Choose the missing product to complete this equation representing the cracking of C11H24?

C11H24 → C9H20 + \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

**C2H4 C3H6 C3H8 C2H6**

1. The cracking of dodecane, C12H26, produces one molecule of propene, one molecule of ethene and a third alkane product.

C12H26 → \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ + C2H4 + C3H6

Choose the molecular formula of the alkane product formed?

**C10H22 C9H20 C8H18 C7H16**

1. The cracking of the alkane C16H34 produces an alkane with 12 carbon atoms in each molecule and a second unsaturated product.

Use some of the formulas provided to complete the equation for the reaction.

**C16H34 C12H24 C12H26 C3H6 C4H8**

C16H34 → \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ + \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

1. Suggest two reasons why catalytic cracking is usually preferred to steam cracking.

*[Hint: Think about the amount of energy needed and what happens to the catalyst.]*

Reason 1: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

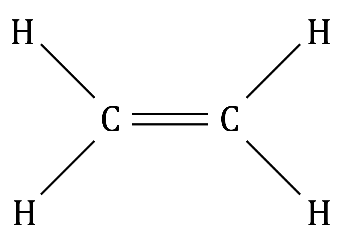
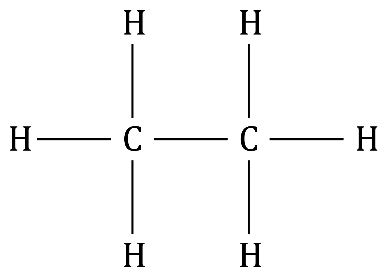
Reason 2: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

**Cracking: feeling confident?**

1. Ethane or ethene? Use some of the displayed structural formulas, words and numbers provided to complete the table comparing ethane and ethene. Some of the words and numbers will not be used.

**0 1 4 5 6 7**

**CH C2H4 C2H6 alkenes alkanes alcohols**



|  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- |
| **Compound** | **Homologous**  **series** | **Molecular formula** | **Displayed structural formula** | **Number of single covalent bonds** | **Number of double covalent bonds** |
| ethane |  |  |  |  |  |
| ethene |  |  |  |  |  |

1. Alkanes that contain between 10 and 16 carbon atoms are the feedstock for steam cracking. By varying the temperature at which cracking is carried out, chemists can control the chain length of the alkanes and alkenes produced.

The table provides information about the alkanes and alkenes produced at two different temperatures during steam cracking.

|  |  |  |  |
| --- | --- | --- | --- |
| **Temperature** | **Point at which the carbon chain breaks** | **Alkanes produced** | **Alkenes produced** |
| higher temperature | near the end of the alkane chain | higher proportion of longer alkanes containing 10 to 14 carbons | higher proportion of smaller alkenes, ethene and propene |
| lower temperature | near the middle of the alkane chain | higher proportion of medium-sized alkanes such as C8H18 | higher proportion of larger alkenes, such as C8H16 |

Use the information provided, and choose from the molecular formulas provided, to write an equation representing the cracking of C14H30 at each of the two temperatures.

*[Hint: You can use the information in the table to decide on the products formed in each process, but the equations must be balanced.]*

**C2H4 C3H6 C7H14 C8H16**

**C13H28 C12H26 C8H18 C9H20 C6H14**

* 1. Higher temperature

C14H30 → \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ + \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

* 1. Lower temperature

C14H30 → \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ + \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

**Cracking: what do I understand?**

Think about your answers and confidence level for each mini-topic. Decide whether you understand it well, are unsure or need more help. Tick the appropriate column.

|  |  |  |  |
| --- | --- | --- | --- |
| **Mini-topic** | **I understand  this well** | **I think I understand this** | **I need more  help** |
| I can compare the supply and demand of fractions produced from the fractional distillation of crude oil. |  |  |  |
| I can describe the reasons for cracking. |  |  |  |
| I can describe the production of smaller alkane molecules and alkene molecules from cracking. |  |  |  |
| I can write equations to  represent cracking. |  |  |  |
| I can compare catalytic and steam cracking. |  |  |  |
| **Feeling confident? topics** | **I understand  this well** | **I think I understand this** | **I need more  help** |
| I can describe the difference between ethane and ethene. |  |  |  |
| I can predict products formed during steam cracking. |  |  |  |