14–16 years

Atomic model

Unscrambling definitions

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Unscrambled definitions

Atomic number is the number of protons in the nucleus of an atom of a particular element.

Mass number is the total number of protons and neutrons in the nucleus of an atom of a particular element.

Relative atomic mass is the average mass of an atom of an element taking into account the naturally occurring percentages of its isotopes.

Relative mass is the mass of a particle relative to $^{1}/_{12}$ of the mass of a ^{12}C atom.

Isotopes are atoms with the same number of protons but different numbers of neutrons.

Relative charge is the positive (+) or negative (-) charge of a particle compared to the charge of a single proton.

Connection completion: completed sentences

Mass number tells us the total number of protons and neutrons in the nucleus of an atom. **Therefore**, the number of neutrons can be calculated **by** subtracting the atomic number from the mass number, **since** atomic number tells us the number of protons in the nucleus of an atom.