Atomic model

Unscramble the phrases in the table to make the correct definitions for the key terms listed in column A (phrases can be used once, more than once, or not at all). Then write out the definitions in full, in the spaces provided below the table.

A	B	C	D	E
Atomic number	is the total number	of protons	of protons but	a particular element.
Mass number	is the average mass of	in the nucleus	and neutrons	a ^{12}C atom.
Relative atomic mass	is the number of protons	with the same number	relative to $^{1}/_{12}$ the mass of	in the nucleus of an atom of a particular element.
Relative mass	are atoms	of a particle	taking into account the naturally	different numbers of neutrons.
Isotopes	is the positive (+) or negative (-) charge	an atom of an element	compared to the charge	occurring percentages of its isotopes.
Relative charge	is the mass	of a particle	of an atom of	of a single proton.

Atomic number _____

_____

Mass number _____

_____

Relative atomic mass _____

_____

Relative mass_____

_____

Isotopes _____

_____

Relative charge _____

_____

Connection completion

Choose the letter from the table below which contains the correct row of connective words to complete these sentences.

Mass number tells us the total number of protons and neutrons in the nucleus of an atom. _____, the number of neutrons can be calculated _____subtracting the atomic number from the mass number, _____ atomic number tells us the number of protons in the nucleus of an atom.

A	Hence	as	nevertheless
B	Since	through	nevertheless
C	Nevertheless	despite	because
D	Therefore	by	since