Structure strips for rates of reaction

Education in Chemistry
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Scaffolding helps students overcome the fear of a blank page. Structure strips provide suitable prompts for a piece of writing, in this case rates of reaction. The student sticks the strip into the margin of their exercise book and writes alongside it.

See the next page.

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Rates structure strip	**Rates structure strip**	**Rates structure strip**	**Rates structure strip**	**Rates structure strip**
Define 'rate of reaction'.	*Define 'rate of reaction'.*	*Define 'rate of reaction'.*	*Define 'rate of reaction'.*	*Define 'rate of reaction'.*
Give four factors that affect the rate of reaction.	*Give four factors that affect the rate of reaction.*	*Give four factors that affect the rate of reaction.*	*Give four factors that affect the rate of reaction.*	*Give four factors that affect the rate of reaction.*
Name the two general ways of measuring rate of reaction.	*Name the two general ways of measuring rate of reaction.*	*Name the two general ways of measuring rate of reaction.*	*Name the two general ways of measuring rate of reaction.*	*Name the two general ways of measuring rate of reaction.*
Explain how collision theory explains rate of reaction.	*Explain how collision theory explains rate of reaction.*	*Explain how collision theory explains rate of reaction.*	*Explain how collision theory explains rate of reaction.*	*Explain how collision theory explains rate of reaction.*
Describe how the following rates of reaction could be determined (include equations and diagrams where necessary). *(i) Mg + HCl* *(ii) $Na_2S_2O_3$ + HCl*	*Describe how the following rates of reaction could be determined (include equations and diagrams where necessary)* *(i) Mg + HCl* *(ii) $Na_2S_2O_3$ + HCl*	*Describe how the following rates of reaction could be determined (include equations and diagrams where necessary)* *(i) Mg + HCl* *(ii) $Na_2S_2O_3$ + HCl*	*Describe how the following rates of reaction could be determined (include equations and diagrams where necessary)* *(i) Mg + HCl* *(ii) $Na_2S_2O_3$ + HCl*	*Describe how the following rates of reaction could be determined (include equations and diagrams where necessary)* *(i) Mg + HCl* *(ii) $Na_2S_2O_3$ + HCl*
Define the term catalyst.	*Define the term catalyst.*	*Define the term catalyst.*	*Define the term catalyst.*	*Define the term catalyst.*
Give some examples of catalysts and the reactions they catalyse.	*Give some examples of catalysts and the reactions they catalyse.*	*Give some examples of catalysts and the reactions they catalyse.*	*Give some examples of catalysts and the reactions they catalyse.*	*Give some examples of catalysts and the reactions they catalyse.*