# Compounds as ions

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[rsc.li/33qaoHE](https://rsc.li/33qaoHE)

Split each compound into its ions:

1. Draw a graphical representation.
2. Write it out.

eg 2NaCl = = 2Na+ + 2Cl-

## Section 1:

NaCl

NaF

LiF

KBr

CaO

BaO

BaS

KI

AlN

GaP

## Section 2:

2NaBr

3KI

4MgO

InN

5SrS

## Section 3:

CaCl2

AlCl3

Na2O

CaO

Al2O3

Li2S

2Li2S

3Li2S

5MgCl2

8Na3N

Potassium fluoride

Potassium oxide

Potassium nitride

Calcium nitride

## Section 4:

NaOH

Na2SO4

Na2CO3

2Na2CO3

6Li2CO3

LiOH

Ca(OH)2

Al(OH)3

4Al(OH)3

MgSO4

2MgSO4

Al2(SO4)3

3Al2(SO4)3

## Section 5:

2K2SO4

CO2

2CH4

4Be(OH)2

Aluminium iodide

Water

Sodium sulfate

## Section 6:

Ga(OH)3 (s)
Ga(OH)3 (l)
Ga(OH)3 (aq)
Cl2(aq)
3NaCl(s)
3NaCl(aq)
A solution of potassium iodide
Melted barium fluoride