# Equilibria

## The equilibrium constant, *K*c

**1.** Match each of the reactions to the correct description;

(2 marks)

Making an ester;

CH3COOH(aq) + CH3OH(aq)

⇌

CH3COOCH3(aq) + H2O(aq)

Heating hydrated

CuSO4.5H2O

Wood burning

An irreversible reaction

A reversible reaction

A reversible reaction at equilibrium

Explain your choice;

(2 marks)

**2.** We can define a constant, *K*c for a reaction at equilibrium. Provided the temperature is constant, the value of *K*c is constant.

For each of the equilibria below, write an expression for the rate constant, *K*c and derive the units of the constant.

1. Electrophilic addition of HCN to propanone;

CH3COCH3(aq) + HCN(aq) ⇌ CH3C(OH)(CN)CH3(aq)

*K*c = Units =

(2 marks)

1. Esterification;

CH3CH2COOH(aq) + CH3OH(aq) ⇌ CH3CH2COOCH3(aq) + H2O(aq)

*K*c = Units =

(2 marks)

1. The Haber Process;

N2(g) + 3 H2(g) ⇌ 2 NH3(g)

*K*c = Units =

(2 marks)

# Calculations with *K*c

**1.** During the Contact process, SO2 is converted into SO3 in a reversible reaction;

2 SO2 + O2 ⇌ 2 SO3 ΔH −197 kJ mol−1

The equilibrium was established at 1000 K and a small sample of the equilibrium mixture extracted. It was found to contain 1.0 mol dm–3 of SO2, 0.2 mol dm–3 of O2 and 1.4 mol dm–3 of SO3.

1. Calculate *K*c at this temperature.

(3 marks)

1. In the Contact process the temperature of choice is 700 K. What effect will this have on the value of *K*c compared to that calculated above?

(1 mark)

**2.** Catherine is studying the hydrolysis of ethyl butanoate;

CH3CH2CH2COOCH2CH3 + H2O ⇌ CH3CH2CH2COOH + CH3CH2OH

She places exactly 1 mol of ethyl butanoate and 2 mol of water in a conical flask and allows the mixture to reach equilibrium. After this time the equilibrium mixture was analysed and found to contain 0.3 mol of butanoic acid. Calculate *K*c for the equilibrium at this temperature.

(3 marks)

**3.** In a different reaction, Catherine wants to make butyl ethanoate. She reacts butanol with ethanoic acid in 50 cm3 of water in a round bottomed flask.

C4H9OH + CH3COOH ⇌ CH3COOC4H9 + H2O

She wishes to make exactly 0.25 mol of butyl ethanoate. If she starts with 0.5 mol of ethanoic acid, how much butanol should she add? (*K*c for the equilibrium at 20 °C is 3.0. The density of water is 1 g cm–3)

(3 marks)

# Le Châtelier and *K*c

Le Châtelier has lost his glasses. He can’t remember which floor of his lab he left them on!

Consider the equilibrium below;

H2(g) + CO2(g) ⇌ H2O(g) + CO(g) ΔH = +40 kJ mol–1

Help Le Châtelier find his glasses by deciding what effect each of the changes in conditions **1**-**9** listed below will have on the value of *K*c for this equilibrium.

Le Châtelier is currently in his office on the second floor

* If the change in conditions increase *K*c, move Le Châtelier one floor up
* If the change in conditions decrease *K*c, move Le Châtelier one floor down
* If the change in conditions have no effect on *K*c, Le Châtelier doesn’t move

A diagram of a floor

Description automatically generatedUnless stated otherwise assume that all conditions other than the one mentioned remain constant.

**1.** Adding a catalyst to the reaction mixture

**2.** Adding CO2 to the reaction mixture

**3.** Increasing the pressure of the system

**4.** Increasing the reaction temperature

**5.** Adding CO to the reaction mixture

**6.** Decreasing the reaction temperature

**7.** Increasing the volume of the reaction container

**8.** Increasing the amount of H2 gas in the reaction mixture

**9.** Increasing the surface area of the catalyst

(9 marks)

Le Châtelier will find his glasses on the floor.

# The equilibrium constant, *K*p

For gases it is easier to measure the pressure of a gas instead of its concentration. Therefore for equilibria involving only gases we quote the equilibrium constant in terms of pressure and give it the symbol *K*p.

**1.** Complete the table below by calculating the equilibrium composition, the mole fractions, the total pressure or the partial pressures for the equilibria shown.

A table of chemical formulas

Description automatically generated(4 marks)

**2.** Write an expression for *K*p, and calculate its value, assuming that each of the systems described above is at equilibrium.

(6 marks)

**The solubility product, *K*sp**

**1.** Define the following terms. Illustrate your definitions with examples from the box.

PbI2(s) ⇌ Pb2+(aq) + 2 I–(aq)

CO(g) + H2O(g) ⇌ CO2(g) + H2(g)

CaCO3(s) ⇌ CaO(s) + CO2(g)

*Homogeneous equilibrium*

*Heterogeneous equilibrium*

(4 marks)

The solubility product is a special example of a heterogeneous equilibrium.

When writing an expression for the equilibrium constant for a heterogeneous equilibrium, terms for pure solids or pure liquids are left out of the expression.

**2.** A student tests for the presence of chloride ions in solution by adding a solution of lead nitrate. A white precipitate of lead chloride is formed.

The precipitate exists in equilibrium with its ions;

PbCl2(s) ⇌ Pb2+(aq) + 2 Cl–(aq), *K*sp = 1.6 × 10–5 mol3 dm–9 at 298 K

1. Write an expression for *K*sp for this equilibrium.

(1 mark)

1. Calculate the solubility of lead chloride at 298 K in mol dm–3.

(2 marks)

1. Describe an experiment the student could do to obtain an experimental value for the solubility product, *K*sp of lead chloride.

(3 marks)

# Equilibria – Answers

## The equilibrium constant, *K*c

1.

A diagram of a reaction

Description automatically generated with medium confidence

*(2 marks for all three correct, 1 mark for 1 correct)*

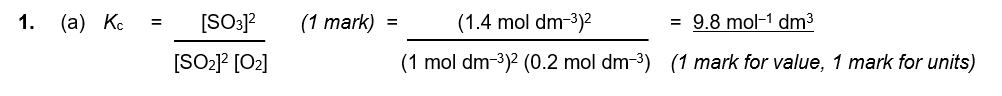
The products from burning wood cannot be turned back into wood so it is irreversible. The copper sulphate once dehydrated can be turned back into the hydrated form by the addition of water. Hence it is a reversible reaction *(1 mark for above two points).* The esterification reaction is in a closed system so neither products nor reactants can escape so it is a reaction at equilibrium *(1 mark).*

2.

A close-up of a math formula

Description automatically generated

**Calculations with *K*c**



(b) The temperature has been decreased. Therefore the equilibrium will shift in favour of the exothermic reaction (to the right) in order to oppose the temperature decrease. Therefore the value of *K*c will increase.

*(1 mark)*

A maths with numbers and formulas

Description automatically generated with medium confidenceA math equations on a white background

Description automatically generatedA white rectangular box with black text

Description automatically generatedA white box with black text

Description automatically generated**Le Châtelier and *K*c**

Le Châtelier can find his glasses on the 2nd floor …… they were in his office all along!

*(1 mark)*

## The equilibrium constant, *K*p

A table with formulas and equations

Description automatically generated with medium confidence*4 marks – one for each box fully completed correctly*

# The solubility product, *K*sp

**1.** *Homogeneous equilibrium* = an equilibrium where all the substances are in the same phase.

e.g. CO(g) + H2O(g) ⇌ CO2(g) + H2(g)

*Heterogeneous equilibrium* = an equilibrium where all the substances are not all in the same phase.

e.g. CaCO3(s) ⇌ CaO(s) + CO2(g) or PbI2(s) ⇌ Pb2+(aq) + 2 I–(aq)

*(2 marks for each correct definition with matching correct example from the table)*

A white paper with black text

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