**Cracking: teacher guidance**

This resource forms part of the **Review my learning** series from the Royal Society of Chemistry. The worksheets assess learner’s understanding of content from common 11–14 and 14–16 curriculums. They can be used to identify knowledge gaps and misconceptions once that part of the curriculum has been taught.

The Cracking worksheets cover the following topics:

* the supply and demand of the different fractions from the fractional distillation of crude oil
* reasons for cracking
* producing smaller alkane molecules and alkenes from cracking
* writing equations to represent cracking
* catalytic and steam cracking.

If learners successfully answer questions on these topics, they can attempt the extension question. This requires learners to compare the formulas and bonding in ethane and ethene, and to use information to write equations for steam cracking reactions at different temperatures.

**Scaffolding**

Level 1 (\*) is a scaffolded worksheet which supports learners in a variety of ways, such as selecting words from a word bank, providing answer options to choose from or completed examples. Level 2 (\*\*) is a partially scaffolded worksheet with a reduced level of support, such as partially completed sentences or a wider range of answer options to select from. Level 3 (\*\*\*) is an unscaffolded worksheet in which most of the tasks involve answering questions with a minimum of prompts.

**Metacognition**

The ‘What do I understand?’ page is common to all levels of worksheet and can be used both to identify areas needing whole class attention and as an indicator for learners to help guide their revision.

Below you will find model answers for each level and guidance on learners’ misconceptions. Learners can use the model answers to self- or peer assess.

**When to use**

The worksheets can be used in a variety of ways:

* To assess learners’ knowledge at the beginning or end of a period of teaching. Match the level of the worksheet to the support needs of the learners.
* To assess knowledge during a period of teaching and after learners have completed the relevant topic.
* As part of revision.
* As a refresher exercise for teachers or non-subject specialists.

There is also scope for the level of the worksheets used to be increased as learners progress through their curriculum.

**Further support**

For more resources to support teaching of this topic and address any misconceptions identified, go to [**rsc.li/3PmRw5E**](https://rsc.li/3PmRw5E).

For more assessment questions on this topic use our Knowledge check and In context worksheets on Cracking hydrocarbons [rsc.li/2SCxbLL](https://rsc.li/2SCxbLL)**.**

**Answers**

Cracking: knowledge check

1. **(Level 1, 2 and 3)**
	1. Demand is higher than supply for **gases**, **petrol**, **kerosene** and **diesel**.
	2. Supply is higher than demand for **bitumen** and **fuel oil**.

**Guidance:** Learners need to understand the meanings of the terms ‘supply’ and ‘demand’.

1. **(Level 1)**
	1. True
	2. False
	3. False
	4. False
	5. True
	6. False

**(Level 2 and 3)**

* 1. True
	2. False – Smaller hydrocarbon molecules are generally more useful than larger hydrocarbon molecules.
	3. False – Unsaturated hydrocarbons contain at least one double covalent bond between carbon atoms.
	4. False – Alkenes are unsaturated compounds.
	5. True
	6. False – Alkanes and alkenes are two different homologous series.

**Guidance:** Common misconceptions include:

* Confusing the process and changes that occur during cracking with the process and changes that occur during fractional distillation. Cracking involves breaking covalent bonds, whereas fractional distillation involves separation by boiling point (which depends on breaking weak intermolecular forces).
* Confusing alkanes and alkenes. Sloppy handwriting may add to this problem as an ‘a’ can easily be confused with an ‘e’. Learners need to understand the terminology, including the meaning of the terms ‘saturated’, ‘unsaturated’ and ‘homologous series’.
1. **(Level 1, 2 and 3)**

demand; alkanes; fuels; diesel; alkenes; polymers

**Guidance:** See guidance for questions 1.1 and 1.2.

1. **(Level 1, 2 and 3)**

covalent; smaller; 550°C; catalyst; speed up; 800°C; steam

**Guidance:** Learners may find ‘steam cracking’ referred to as ‘thermal cracking’ in other sources.

Cracking: test myself

1. **(Level 1, 2 and 3)**

thermal decomposition/decomposition

**Guidance:** Learners need to be familiar with the meaning of the term ‘decomposition’ in relation to chemical changes.

1. **(Level 1, 2 and 3)**

C*n*H2*n*

**Guidance:** Misconceptions include:

* not understanding the nature of a general molecular formula
* confusing general molecular formulas with molecular formulas and empirical formulas.
1. **(Level 1, 2 and 3)**

two pairs of electrons

**Guidance:** A common misconception is confusing the number of pairs of electrons shared with the number of electrons shared.

1. **(Level 1, 2 and 3)**

ethene

**Guidance:** Learners may confuse the meanings of the terms ‘saturated’ and ‘unsaturated’.

1. **(Level 1)**

C2H4

**(Level 2 and 3)**

C11H24 → **C9H20** + **C2H4**

**Guidance:**Learners can check accuracy by adding up the numbers of ‘H’s and ‘C’s on either side of the equation.

Common errors include not using upper case letters and subscripts correctly in formulas and not balancing the equation correctly.

1. **(Level 1, 2 and 3)**

C7H16

**Guidance:** Misconceptions include:

* not correctly calculating the number of carbon atoms in the alkane produced
* not using the correct molecular formula for heptane.
1. **(Level 1, 2 and 3)**

C16H34 → **C12H26** +**C4H8**

**Guidance:** Misconceptions include:

* not correctly calculating the number of carbon atoms in the alkene produced
* writing the incorrect molecular formula for the alkene.
1. **(Level 1, 2 and 3)**

Reason 1: **it uses less energy.**

Reason 2:  **the catalyst can be reused.**

**Guidance:** This is a synoptic question and requires learners to link cracking methods to climate change strategies and the advantages of using catalysts.

Cracking: feeling confident?

1. **(Level 1, 2 and 3)**

|  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- |
| **Compound** | **Homologousseries** | **Molecular formula** | **Displayed structural formula** | **Number of single covalent bonds** | **Number of double covalent bonds** |
| ethane | **alkanes** | **C2H6** | Two letters C in the middle of the diagram joined by a single line representing a single bond. Each letters C has three letters H joined by single lines to represent single bonds.  | **7** | **0** |
| ethene | **alkenes** | **C2H4** | Two letters C joined with two lines representing a double bond and each letter C has two letters H joined with a single line representing a single bond. | **4** | **1** |

**Guidance:** Misunderstandings arise from not recognising the notation used to represent single and double covalent bonds.

1. **(Level 1, 2 and 3)**

Learners’ answers may vary but examples include:

* 1. Higher temperature

C14H30 → **C12H26** + **C2H4**

* 1. Lower temperature

C14H30 → **C6H14** + **C8H16**

**Guidance:** There are alternative answers to this question. The products from cracking are not exact or predictable, and several different reactions are possible. Completing this question may help learners to appreciate that.

**Cracking: what do I understand?**

|  |  |
| --- | --- |
| **Mini-topic** | **Assessed via:** |
| I can compare the supply and demand of fractions produced from the fractional distillation of crude oil. | 1.1 |
| I can describe the reasons for cracking. | 1.2, 1.3 |
| I can describe the production of smaller alkane molecules and alkene molecules from cracking. | 1.4, 2.1, 2.2, 2.3 |
| I can write equations torepresent cracking. | 2.4, 2.5, 2.6, 2.7 |
| I can compare catalytic and steam cracking. | 2.8 |
| **Feeling confident? topics** | **Assessed via:** |
| I can describe the difference between ethane and ethene. | 3.1 |
| I can predict products formed during steam cracking. | 3.2 |