

## Explaining observations in energetics quiz: teacher notes and answers

### *Education in Chemistry*

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This quiz follows the pointers in the *EiC* article 'How to make quizzes more worthwhile'. It is aimed at the 11–14 age group.

1. When zinc powder is added to sulfuric acid the temperature increases quickly and then slowly decreases.  
Which of the following provides the best explanation for what is taking place?
  - A. The reaction between zinc and sulfuric acid is exothermic.
  - B. The temperature rises as zinc and sulfuric acid react, and then decreases back to room temperature as the heat energy is lost to the surroundings.
  - C. The reaction between zinc and sulfuric acid is exothermic so the temperature in the test tube rises.
  - D. The temperature rises quickly as zinc and sulfuric acid react exothermically but decreases as heat energy is slowly transferred to the mixture of air particles.
  
2. When copper is added to hydrochloric acid, the temperature of the acid does not change.  
Which of the following explanations for this observation is least helpful?
  - A. Copper is a metal.
  - B. Copper is less reactive than hydrogen so it does not react with the acid.
  - C. No energy transfer is taking place.
  - D. There is no exothermic or endothermic change taking place.

For questions 3–5, use the following table to choose whether the answer is A, B, C, or D:

A	B	C	D
I and III	II and IV	I, II, and III	IV only

3. When 0.3 g of magnesium ribbon is added to 20 cm<sup>3</sup> of hydrochloric acid (which is in excess), the temperature increases by 22°C. Which (combination) of the following would decrease the maximum temperature reached?

- I. Using 0.3 g of magnesium powder.
- II. Using 50 cm<sup>3</sup> of hydrochloric acid.
- III. Using more concentrated hydrochloric acid.
- IV. Using 0.1 g of magnesium ribbon.

B

4. Which (combination) of the following would increase the temperature rise of the water in the conical flask?

- I. Use half the volume of water in the conical flask.
- II. Moving the conical flask down so the flame turns yellow-orange and leaves soot on the base of the flask.
- III. Using a longer chain alcohol.
- IV. Moving the conical flask to the top of the clamp stand.

A

5. When the white solids barium hydroxide and ammonium hydroxide are mixed in a beaker a chemical reaction takes place. If the beaker is placed on a heat proof mat with a few drops of water on it, the mat will stick to the bottom of the beaker.

Which (combination of) statement(s) explains why the mat sticks to the bottom of the beaker?

- I. The water molecules have less energy so the water freezes.
- II. The barium hydroxide and ammonium chloride undergo an endothermic reaction in the beaker.
- III. Heat energy is transferred from the surroundings into the beaker.
- IV. Water is formed in the reaction.

C

