

Elements, compounds and mixtures

Unscramble the phrases in the table to make the correct definitions for the key terms listed in column A (phrases can be used once, more than once, or not at all). Then write out the definitions in full, in the spaces provided below the table.

A	B	C	D	E
The molecular formula	is when a substance is	a lot of		
The empirical formula	uses chemical symbols to show the	dissolved in	each element in one molecule of the substance,	shown by the state symbol (aq).
The chemical formula	is when no more solute can	not very much	each element in a substance,	that temperature.
Aqueous	uses chemical symbols to give the	dissolve in a particular		such as N ₂ for nitrogen or H ₂ O for water.
Dilute	gives the	relative number of the atoms of	solute dissolved in a	such as CH ₂ for ethene which has molecular formula C ₂ H ₄ .
Concentrated	is when there is	number of atoms of	volume of solvent at	such as H ₂ O for water or NaCl for sodium chloride.
A saturated Solution		simplest whole number ratio of atoms of	water;	particular volume of solvent.

The molecular formula _____

The empirical formula _____

The chemical formula _____

Aqueous _____

Dilute _____

Concentrated _____

A saturated solution _____

Connection completion

Choose the letter from the table below that contains the correct row of connective words to complete these sentences:

A dilute, aqueous solution can be made more concentrated _____ adding more solute. _____, in order to produce a saturated solution, more solute must be added _____ no more can dissolve in a particular volume of solvent at that temperature.

A	despite	Although	or
B	through	Despite this	therefore
C	by	Furthermore	until
D	thereby	Conversely	in addition